



General Certificate of Secondary Education
2024

Centre Number

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Candidate Number

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Single Award Science

Unit 4

Booklet A

Foundation Tier

ML

[GSA41]

TIME

2 hours, plus your additional time allowance.

INSTRUCTIONS TO CANDIDATES

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.

Write your answers in the spaces provided in this question paper.
Answer **both** questions.

INFORMATION FOR CANDIDATES

The total mark for this paper is **30**.

Figures in brackets printed down the right-hand side of pages indicate the marks awarded to each question or part question.

Follow all health and safety instructions.

You may use a ruler and calculator if required.

The apparatus and materials required to complete the task(s) are provided.

A Data Leaflet, which includes a Periodic Table of the Elements, is included in this question paper.

For Examiner's
use only

Question Number	Marks
1	
2	

Total
Marks

Task 1 – Biology

Examiner Only	
Marks	Remark

Investigating photosynthesis

1 You have been given a destarched plant for this investigation.

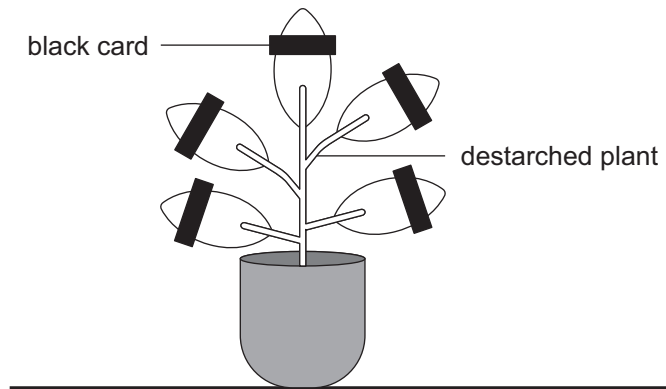
- (a) Complete the sentence below to describe how the plant was destarched.

Choose from:

light dark cold two forty-eight

The plant was destarched by leaving it in a _____
place for _____ hours. [2]

The plant was destarched. Then a strip of black card was placed over part of each leaf. You can see this in the diagram below.



The plant was then left in bright light for a period of time.

- (b) Name the factor needed for photosynthesis which is missing from the part of the leaf under the black card.

_____ [1]

Examiner Only	
Marks	Remark

- (c) (i) Table 1**

[1]

(ii) Table 2

[2]

Circle your answer.

[1]

- (e) Why was a Bunsen burner **not** used to heat the boiling tube containing ethanol?

_____ [1]

- (f) In **step 6**, the ethanol was used to remove the chlorophyll from the leaf. Suggest why it was necessary to remove the chlorophyll.

_____ [1]

- (g) Explain why, in **step 9**, the leaf was placed back into the hot water after removing it from the beaker containing ethanol.

Circle your answer.

to soften it

to stop chemical reactions

to kill the leaf

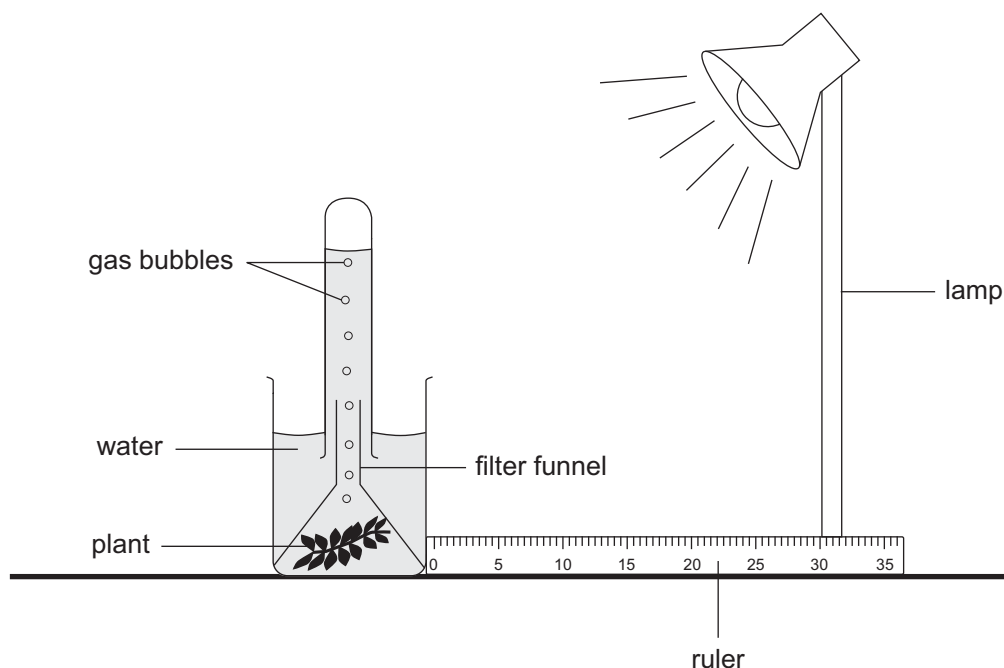
[1]

Examiner Only	
Marks	Remark

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(Questions continue overleaf)

- (h) Another investigation was carried out to show how light intensity affects the **rate** of photosynthesis.

The diagram below shows how the apparatus was set up.



The lamp was placed 10 cm from the plant at the start of the investigation and the number of bubbles of gas produced in one minute was counted.

The lamp was then moved different distances from the plant.

The number of bubbles of gas produced in one minute at each distance was counted.

The table below shows the results.

Distance of lamp from the plant/cm	Number of bubbles of gas produced in one minute
10	56
15	44
20	31
25	19
30	7

Task 2 – Chemistry

Examiner Only	
Marks	Remark

Investigating neutralisation

- 2 Indigestion tablets contain an alkali and work by neutralising excess acid in the stomach.

You are provided with a beaker of acid and one crushed indigestion tablet.

- (a) (i) Name the hazard symbol shown on the acid.

Choose from:

corrosive **caution** **explosive** **flammable**

_____ [1]

- (ii) State **one** safety precaution needed when using the acid.

_____ [1]

- (b) Describe the appearance of the crushed indigestion tablet.

_____ [1]

- (c) (i) Follow the method below.

1. Use the measuring cylinder to measure 20 cm^3 of water and pour this into a small beaker.
2. Add the crushed indigestion tablet to the beaker and stir until dissolved.
3. Add 2–3 drops of universal indicator solution to the beaker.
4. Record the colour of the solution **before** any acid is added in the table opposite.
5. Use the pH chart to find the pH of the solution and record this in the table.
6. Use the dropping pipette to add 1 cm^3 of hydrochloric acid to the solution in the beaker.
7. Stir the solution and record its colour and pH in the table.
8. Repeat steps **6** and **7** until you have added a total of 10 cm^3 of acid to the beaker.

Results table

Volume of acid added/cm ³	Colour of universal indicator	pH
0		
1		
2		
3		
4		
5		
6		
7		
8		
9		
10		

[3]

- (ii) Suggest the volume of acid added that made the solution pH 7.

_____ cm³ [1]

- (d)** State **one** variable that was kept constant during the investigation to make sure the results were valid (fair test).

[1]

- (e)** Name the piece of apparatus that could have been used to measure a more accurate pH value.

Circle your answer.

thermometer : measuring cylinder : pH meter [1]

Examiner Only	
Marks	Remark

(ii) State the trend shown by these results.

[1]

(iii) Use the results opposite to suggest the volume of acid that will be needed to neutralise **five** indigestion tablets.

_____ cm³ [1]

(g) The teacher suggested that the student should repeat his investigation and average the results. Why is this important?

Circle your answer.

to make the results more accurate

to make the results more precise

to make the results more reliable

[1]

THIS IS THE END OF THE QUESTION PAPER

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